**Skills Portfolio Name: …………………………….**

**Chemistry Practical 3: Standard solutions and titrations**

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| Making up a standard solution Level of confidence (circle) 1 (low) 2 3 4 5 (high) | |
| *Insert photograph here* | ***What was difficult about the technique? What advice would you give another student to carry it out correctly?*** |

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| Calculating concentration Level of confidence (circle) 1 (low) 2 3 4 5 (high) |
| ***Calculate the concentration of your standard solution based on the mass of sodium carbonate you used. Show all working.*** |

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| Use of a volumetric pipette Level of confidence (circle) 1 (low) 2 3 4 5 (high) | |
| *Insert photograph here* | ***What was difficult about the technique? What advice would you give another student to carry it out correctly?*** |

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| Filling a burette Level of confidence (circle) 1 (low) 2 3 4 5 (high) | |
| *Insert photograph here* | ***What was difficult about the technique? What advice would you give another student to carry it out correctly?*** |

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| Performing a titration Level of confidence (circle) 1 (low) 2 3 4 5 (high) | |
| *Insert photograph here* | ***What was difficult about the technique? What advice would you give another student to carry it out correctly?*** |

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| Titration calculations (1) Level of confidence (circle) 1 (low) 2 3 4 5 (high) |
| ***Calculate the concentration of sodium hydroxide based on your results from part A of the practical. Show all working.*** |

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| Titration calculations (2) Level of confidence (circle) 1 (low) 2 3 4 5 (high) |
| ***Calculate the concentration of your partner group’s solution of sodium carbonate based on your results from part B of the practical. Show all working.*** |